

Chemistry C1 and C2 – Atomic structure and Periodic table

Key Knowledge Questions

1. What are all substances made up of?
2. What do the following symbols stand for (s), (l), (g)?
3. What is the name of two or more substances that **are not** chemically combined together?
4. What is the name of two or more substances that **are** chemically combined together?
5. What is the name for a material that is made up of only 1 type of atom?
6. What process is used to separate liquids with different boiling points?
7. What process is used to separate a liquid and an insoluble solid?
8. What process is used to detect the different dyes in food colourings?
9. Who came up with the idea of the 'Plum Pudding' model?
10. Which Scientist proved that there was a dense positive nucleus?
11. What is the name of the negative subatomic Particle?
12. What is the name of the positive subatomic Particle?
13. What is the name of the neutral subatomic Particle?
14. What does the Atomic Number of an atom show?
15. What is the atomic mass number the sum of?
16. What do you call an atom of the same element with different numbers of neutrons?
17. What is the name of a charged atom?
18. What is the name of group 0?
19. What is the name of group 1?
20. What is the name of group 7?
21. What gas is formed when an alkali metal reacts with water?
22. How does the reactivity of the halogens change as you go up the group?
23. How does the reactivity of the alkali metals change as you go down the group?
24. How many electrons are on the outer shell of group 1 metals?
25. How many electrons are on the outer shell of group 7 elements?
26. Are the metals on the left or the right of the periodic table?
27. Are the non-metals on the left or the right of the periodic table?

Key Knowledge Answers

1. atoms
2. solid, liquid and gas
3. mixture
4. compound
5. element
6. fractional distillation
7. filtering
8. chromatography
9. JJ Thompson
10. Ernest Rutherford
11. electron
12. proton
13. Neutron
14. The number of protons
15. The number of protons and neutrons
16. Isotope
17. Ion
18. Nobel gasses
19. Alkali metals
20. Halogens
21. Hydrogen
22. They become more reactive
23. They become more reactive
24. 1 electron
25. 7 electrons
26. Left
27. right